

Sem 2: Chemistry Core Course

Quantum Numbers

What are Quantum Numbers?

The values of the conserved quantities of a quantum system are given by quantum numbers. Electronic quantum numbers (the quantum numbers describing electrons) can be defined as a *group of numerical values which provide solutions that are acceptable by the Schrodinger wave equation for hydrogen atoms.*

Four quantum numbers can be used to completely describe all the attributes of a given electron belonging to an atom, these are:

- Principal quantum number, denoted by n .
- Orbital angular momentum quantum number (or azimuthal quantum number), denoted by l .
- Magnetic quantum number, denoted by m_l .
- The electron spin quantum number, denoted by m_s .

Number	Symbol	Possible Values
Principal Quantum Number	n	1, 2, 3, 4, ...
Angular Momentum Quantum Number	l	0, 1, 2, 3, ..., $(n - 1)$
Magnetic Quantum Number	m_l	$-l, \dots, -1, 0, 1, \dots, l$
Spin Quantum Number	m_s	$+1/2, -1/2$

The Four Quantum Numbers that Describe an Electron

Principal Quantum Number

- *Principal quantum numbers are denoted by the symbol 'n'. They designate the principal electron shell of the atom.* Since the most probable distance between the nucleus and the electrons is described by it, a larger value of the principal quantum number implies a greater distance between the electron and the nucleus (which, in turn, implies a greater atomic size).
- The value of the principal quantum number can be any integer with a positive value that is equal to or greater than one. The value $n=1$ denotes the innermost electron shell of an atom, which corresponds to the lowest energy state (or the ground state) of an electron.

- Thus, it can be understood that the principal quantum number, n , cannot have a negative value or be equal to zero because it is not possible for an atom to have a negative value or no value for a principal shell.
- When a given electron is infused with energy (excited state), it can be observed that the electron jumps from one principle shell to a higher shell, causing an increase in the value of n . Similarly, when electrons lose energy, they jump back into lower shells and the value of n also decreases.
- The increase in the value of n for an electron is called absorption, emphasizing the photons or energy being absorbed by the electron. Similarly, the decrease in the value of n for an electron is called emission, where the electrons emit their energy.

Azimuthal Quantum Number (Orbital Angular Momentum Quantum Number)

- The azimuthal (or orbital angular momentum) quantum number describes the shape of a given orbital. It is denoted by the symbol ' l ' and its value is equal to the total number of angular nodes in the orbital.
- A value of the azimuthal quantum number can indicate either an s, p, d, or f subshell which vary in shapes. This value depends on (and is capped by) the value of the principal quantum number, i.e. the value of the azimuthal quantum number ranges between 0 and $(n-1)$.
- For example, if $n = 3$, the azimuthal quantum number can take on the following values - 0, 1, and 2. When $l=0$, the resulting subshell is an 's' subshell. Similarly, when $l=1$ and $l=2$, the resulting subshells are 'p' and 'd' subshells (respectively). Therefore, when $n=3$, the three possible subshells are 3s, 3p, and 3d.
- In another example where the value of n is 5, the possible values of l are 0, 1, 2, 3, and 4. If $l = 3$, then there are a total of three angular nodes in the atom.

The allowed subshells under different combinations of ' n ' and ' l ' are listed below. It can be understood that the '2d' orbital cannot exist since the value of ' l ' is always less than that of ' n '.

Magnetic Quantum Number

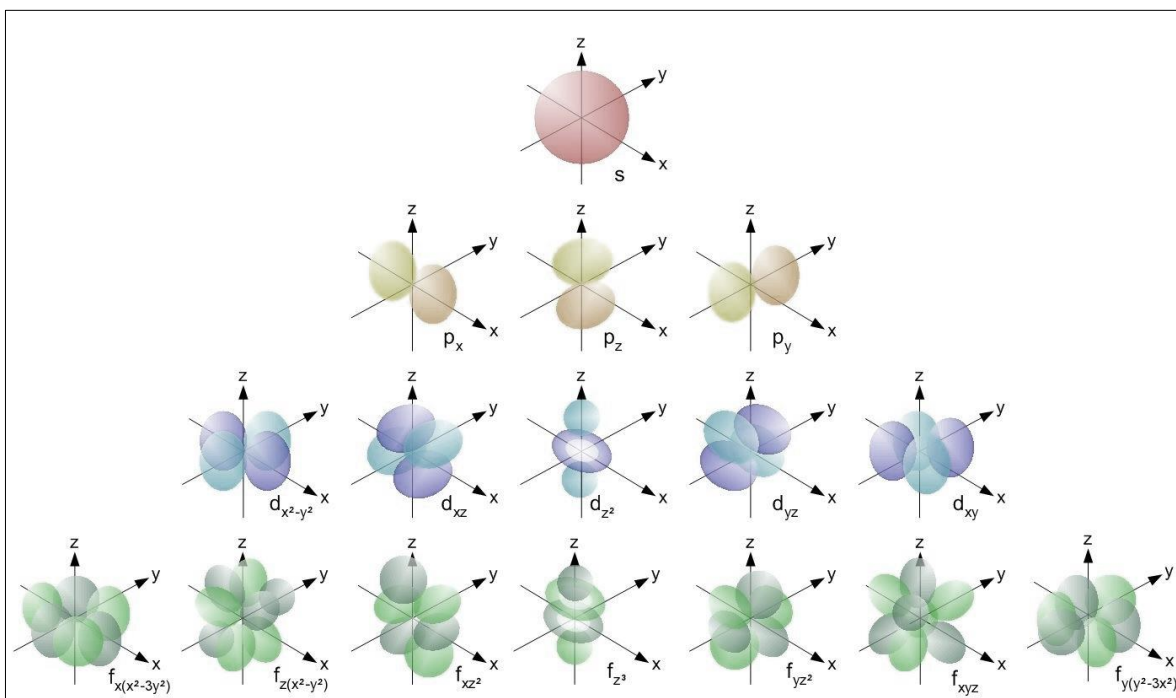
The total number of orbitals in a subshell and the orientation of these orbitals are determined by the magnetic quantum number. It is denoted by the symbol ' m_l '. This number yields the projection of the angular momentum corresponding to the orbital along a given axis.

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	$l=0$	$l=1$	$l=2$	$l=3$	$l=4$	$l=5$
$n=1$	1s	Forbidden combinations of quantum numbers				
$n=2$	2s	2p	Forbidden combinations of quantum numbers			
$n=3$	3s	3p	3d	Forbidden combinations of quantum numbers		
$n=4$	4s	4p	4d	4f	Forbidden combinations of quantum numbers	
$n=5$	5s	5p	5d	5f	5g	Forbidden combinations of quantum numbers
$n=6$	6s	6p	6d	6f	6g	6h
$n=7$	7s	7p	7d	7f	7g	7h ...
$n=8$	8s	8p	8d	8f	8g	8h ...

Allowed Combinations of quantum numbers

Combinations of the Principal and Azimuthal Quantum Numbers



Shapes of Orbitals (as per the corresponding Quantum Numbers)

The value of the magnetic quantum number is dependent on the value of the azimuthal (or orbital angular momentum) quantum number. For a given value of l , the value of m_l ranges between the interval $-l$ to $+l$. Therefore, it indirectly depends on the value of n .

For example, if $n = 4$ and $l = 3$ in an atom, the possible values of the magnetic quantum number are $-3, -2, -1, 0, +1, +2$, and $+3$.

Azimuthal Quantum Number Value	Corresponding Number of Orbitals ($2l + 1$)	Possible Values of m_l
0 ('s' subshell)	$2 \cdot 0 + 1 = 1$	0
1 ('p' subshell)	$2 \cdot 1 + 1 = 3$	$-1, 0$, and 1
2 ('d' subshell)	$2 \cdot 2 + 1 = 5$	$-2, -1, 0, 1$, and 2
3 ('f' subshell)	$2 \cdot 3 + 1 = 7$	$-3, -2, -1, 0, 1, 2$, and 3

The total number of orbitals in a given subshell is a function of the ' l ' value of that orbital. It is given by the formula $(2l + 1)$. For example, the '3d' subshell ($n=3, l=2$) contains 5 orbitals ($2 \cdot 2 + 1$). Each orbital can accommodate 2 electrons. Therefore, the 3d subshell can hold a total of 10 electrons.

Electron Spin Quantum Number

- The electron spin quantum number is independent of the values of n, l , and m_l . The value of this number gives insight into the direction in which the electron is spinning, and is denoted by the symbol m_s .
- The value of m_s offers insight into the direction in which the electron is spinning. The possible values of the electron spin quantum number are $+\frac{1}{2}$ and $-\frac{1}{2}$.
- The positive value of m_s implies an upward spin on the electron which is also called 'spin up' and is denoted by the symbol \uparrow . If m_s has a negative value, the electron in question is said to have a downward spin, or a 'spin down', which is given by the symbol \downarrow .
- The value of the electron spin quantum number determines whether the atom in question has the ability to produce a magnetic field. The value of m_s can be generalized to $\pm\frac{1}{2}$.

Summary

In order to simplify the details of the four different quantum numbers that are related to atomic physics, a tabular column detailing their names, symbols, meanings, and possible values is provided below.

Name and Symbol	Meaning and Possible Values
Principal quantum number, n	Electron shell, $n \geq 1$
Azimuthal quantum number, l	Subshells ($s=0, p=1, \text{etc.}$), $(n-1) \geq l \geq 0$
Magnetic quantum number, m_l	Total number and orientation of orbitals, $ \leq m_l \leq l$
Electron spin quantum number, m_s	The direction of electron spin, $m_s = \pm \frac{1}{2}$

It is important to note that it is impossible for two electrons of the same atom to have exactly the same quantum state or exactly the same values of the set of quantum numbers.

Solved Examples

What are the Possible Subshells when $n = 4$? How Many Orbitals are Contained by Each of these Subshells?

When $n = 4$, the possible l values are 0, 1, 2, and 3. This implies that the 4 possible subshells are the 4s, 4p, 4d, and 4f subshells.

- The 4s subshell contains 1 orbital and can hold up to 2 electrons.
- The 4p subshell contains 3 orbitals and can hold up to 6 electrons.
- The 4d subshell contains 5 orbitals and can hold up to 10 electrons.
- The 4f subshell has 7 orbitals and can hold up to 14 electrons.

Thus, a total of 4 subshells are possible for $n = 4$.

What are the Possible m_l values for $l = 4$?

Since the value of the magnetic quantum number ranges from $-l$ to l , the possible values of m_l when $l = 4$ are: -4, -3, -2, -1, 0, 1, 2, 3, and 4.

Ref: (i) <https://byjus.com/chemistry/quantum-numbers/>

(ii) Inorganic Chemistry, A.K. Das. Volume 1

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